

Acids, bases and salts

Mark scheme

1.	(a)	water / H ₂ O / hydrogen oxide	1	
	(b)	eg H (atom) loses an electron to form H ⁺ or <u>only</u> a proton left	1	
	(c)	is partially ionised in water	1	
	(d)	(i) eg same concentration / quantity of Mg <i>accept: volume of acid / ribbon for both / same time</i> <i>accept: volume of gas measured under the same conditions</i>	1	
		(ii) C A D B	1	
	(e)	(i) OH ⁻	1	
		(ii) acidic	1	[7]
2.	(a)	(i) water <i>accept H₂O</i> <i>accept correct ringed answer in box</i>	1	
		(ii) neutralisation <i>accept underlining or any indication, eg tick</i>	1	
	(b)	sodium hydroxide	1	
		sulphuric acid <i>apply list principle total</i>	1	[4]
3.	(a)	sodium ions and chloride ions (not chlorine) <i>allow sodium chloride/salt/common salt</i> <i>for 1 mark</i>	1	
	(b)	H ⁺ + OH ⁻ → H ₂ O H ⁺ from (hydrochloric) acid OH ⁻ from alkali/sodium hydroxide lose 1 mark if no charge shown disregard other ions <i>each for 1 mark</i>	3	[4]